



Double Replacement Reactions

Lab #10

When two solutions containing dissolved ionic compounds are mixed with each other, they may or may not react with one another. If no visible change is noted, then it is likely that no reaction has taken place. If a precipitate (solid) or a gas is produced, then a reaction has occurred. Through this lab, we will perform numerous combinations of solutions and determine the combinations that produce reactions and those that do not.

Materials:

Well plate

Solutions containing:

NaCl	KI	Zn(C ₂ H ₃ O ₂) ₂	Na ₂ CrO ₄
Cu(C ₂ H ₃ O ₂) ₂	H ₂ SO ₄	Na ₂ CO ₃	Ca(NO ₃) ₂
Mg(NO ₃) ₂	NaOH	MgSO ₄	Pb(NO ₃) ₂
K ₂ SO ₄	Ba(NO ₃) ₂	K ₃ PO ₄	FeCl ₃

Procedure:

Place approximately ten drops of one solution into one of the wells of the well plate. Add ten drops of another solution to the same well and observe the results. Repeat this procedure until all possible combinations of solutions have been reacted.

Results:

Write balanced chemical equations for each combination of solutions that produced a reaction.

LMHSC Labs

Lab Title (cont.)

Lab #